COURSE DESCRIPTION CARD - SYLLABUS

Course name

Influence of Electromagnetic Radiation on Matter

Course

Field of study	Year/Semester
Chemical Technology	II/4
Area of study (specialization)	Profile of study
-	general academic
Level of study	Course offered in
First-cycle studies	English
Form of study	Requirements
full-time	elective

Number of hours

Lecture	Laboratory classes	Other (e.g. online)
0	15	0
Tutorials	Projects/seminars	
0	0	

Number of credit points

2

Lecturers

Responsible for the course/lecturer:

Prof. Andrzej Lewandowski

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Wydział Technologii Chemicznej

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Responsible for the course/lecturer:

Prerequisites

Students:

have knowledge in general chemistry (writing chemical reactions, converting concentrations, knowledge of laboratory glassware and basic laboratory equipment).

have knowledge in mathematics and physics enabling the introduction of problems in physical chemistry (basic laws of physics, differential calculus).

are able to prepare solutions of specific concentrations.

are aware of further development of their competences.

Course objective

To familiarise students with basic problems in physical chemistry at the academic level in the field of: nuclear chemistry, properties of molecules and spectroscopic methods.

Course-related learning outcomes

Knowledge

Students will be able to define and explain selected problems in nuclear chemistry (properties of ionising radiation). K_W03, K_W10

Students will be able to characterise chemical compounds using selected spectroscopic methods. K_W03, K_W11

Skills

Students will be able to obtain information from literature, databases and other sources; interpret it as well as draw conclusions and formulate and substantiate opinions. K_U01

Students will be able to work individually and as part of a team; estimate the time needed to complete the assigned task. K_U2

Students will be able to apply the principles of thermodynamics in the implementation of chemical processes. K_U23

Students will have the self-study skills in the subject. K_U05

Students will be able to elaborate, describe and present results of an experiment or theoretical calculations. K_U09

Students will be able to distinguish between types of chemical reactions and to select them for specific chemical processes. K_U18

Social competences

Students will understand the need for further training and developing their professional competences. K_K01

Students will be able to properly prioritise the task. K_K04

Methods for verifying learning outcomes and assessment criteria

Learning outcomes presented above are verified as follows:

Laboratory classes: The course passing is based on points received for the individual exercise description. Passing exercises from 56% .

Programme content

Laboratory classes:

Dipole moment: permanent and induced. Dielectric constant and capacity of capacitor. Orientation polarization, induced polarization, polarizability, refraction. Additivity of refraction. Ion susceptibility to deformation. Polarizabilities and molecular structures. Influence of molecule structure on the value of its dipole moment. Dipole moment of different types of bonds. Refractometria.

Electromagnetic radiation. Absorption of radiation by matter. Mechanisms of excited system transition to ground state. The spectrum of radiation with particular regard to the visible range (Vis). Seeing colors (color and its complement). Color mixing. Dyes-colored organic and inorganic compounds, their structure. PH indicators - color change mechanism. Spectrophotometer construction. Principles of the spectrophotometric measurements. Lambert-Beer law.

Nuclear nucleus, components, energy, nuclear forces. Elemental particles. Nuclear transformations. Properties of ionizing radiation. Interaction of ionizing radiation with matter. Radiometry and dosimetry. Ionizing radiation detectors. Types and use of radiation sources - open and closed sources. Principles of radiation protection.

Teaching methods

Laboratory classes- practical method - laboratory exercises. Planning, execution and analysis of the results of physicochemical experiment.

Bibliography

Basic

1. P. Atkins, Physical Chemistry, Oxford University Press

2. RS. Barry, SA. Rice, J. Ross, Physical Chemistry, Wiley & Sons, New York 1980.

Additional

1. Naftaly Menn: Practical optics. Elsevier, 2004, s. 193–195

2. Physical Chemistry Instructions: http://zchf.fct.put.poznan.pl.

3. Thermodynamics Lab Instructions http://moodle.put.poznan.pl

Breakdown of average student's workload

	Hours	ECTS
Total workload	45	2,0
Classes requiring direct contact with the teacher	25	1,1
Student's own work (preparation for laboratory classes, preparation of the report.) [1]	20	0,9

[1] delete or add other activities as appropriate